Marking Scheme for WBJEE 2026

Understanding the marking scheme is essential for efficient answering and attempting questions with a strategic mindset. Check marking scheme in the table provided below:

WBJEE Marking Scheme 2026		
Category	**Marks Distribution**	**Negative Marking**
Category I	1 mark	¼ mark
Category II	2 mark	½ mark
Category III	2 mark	No Negative Marking

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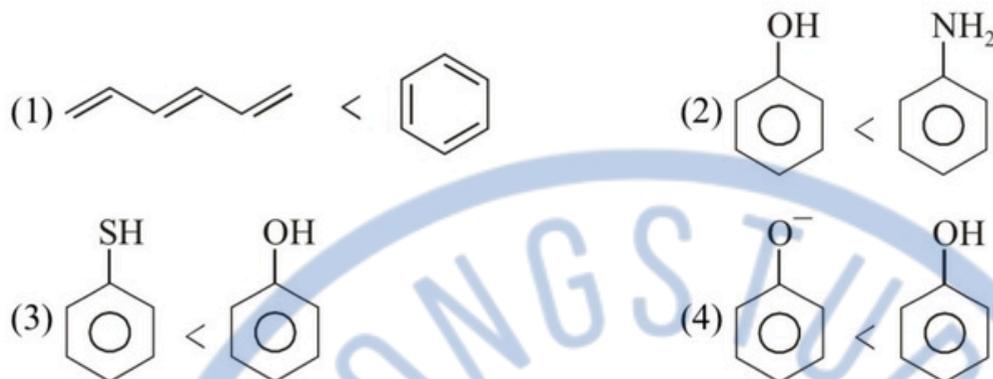
CHEMISTRY

Category-1

1. There is 34 g H_2O_2 in its 2L aqueous solution. The volume strength of the solution is :
   (1) 33.6 V
   (2) 5.6 V
   (3) 22.4 V
   (4) 11.2 V

2. Syn gas is a mixture of :
   (1) $CO_2 + H_2$
   (2) $CO + H_2$
   (3) $CO + CO_2$
   (4) $CO + N_2$

3. Which order is incorrect according to their resonance energy :
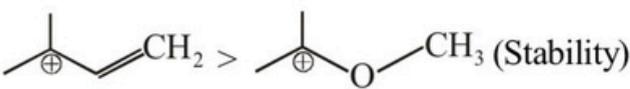
   (1) <

   (2) <

   (3) <

   (4) <

4. The ratio of mass percent of C and H of an organic compound $C_xH_yO_z$ is 7.5 : 1. If the percentage of oxygen in the compound is 32%. Then, the empirical formula of the compound is :
   (1) C_2H_4O
   (2) C_5H_4O
   (3) $C_3H_3O_2$
   (4) $C_5H_8O_2$

5. The correct order in which the O–O bond length increases in the following is :
   (1) $O_3 < H_2O_2 < O_2$
   (2) $O_2 < O_3 < H_2O_2$
   (3) $O_2 < H_2O_2 < O_3$
   (4) $H_2O_2 < O_2 < O_3$

6. Which of the following group has the maximum hyperconjugation effect when attached to benzene ring?
   (1) $(CH_3)_3C-$
   (2) CH_3-CH_3-
   (3) $(CH_3)_2CH-$
   (4) CH_3-

7. 3L mixture of $O_2(g)$ and $N_2(g)$ is passed through ozonizer. The gas mixture obtained contains O_2, O_3 and N_2 gases only. This obtained mixture is passed through alkaline pyrogallol and then through turpentine oil, there were 1100 mL and 600 mL volume contractions respectively. The volume of $O_2(g)$ and $N_2(g)$ in the original mixture respectively is -
   (1) 1.5 L, 1.5 L
   (2) 2.5 L, 0.5 L
   (3) 1L, 2 L
   (4) 2 L, 1 L

8. In the solvay process of manufacturing of Sodium Bicarbonate, the final by-product is :
(1) NH_4Cl (2) $NaHCO_3$ (3) $CaCl_2$ (4) CO_2

9. Which comparison is not correct as indicated ?

(1) $C_6H_5-OH > CH_3OH$ (Acidic nature)

(2) ![structure] $>$![structure] (Stability)

(3) ![pyridine] $>$![pyrrole] (Basic nature)
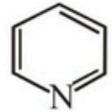
(4) $R-\overset{O}{\underset{O}{\overset{\|}{\underset{\|}{S}}}}-\overset{\ominus}{O} > R-\overset{O}{\overset{\|}{C}}-O^{\ominus}$ (Stability)

10. 15 g of a gas present in 5L vessel exerts a pressure of 0.1 bar. What is the root mean square speed of the gas molecules?
(1) 50 m/s   (2) 10^2 cm/s   (3) 100 m/s   (4) 75 m/s

11. In which of the following pair both the chlorides do not impart colour to the flame?
(1) $BeCl_2$ and $SrCl_2$ (2) $BeCl_2$ and $MgCl_2$
(3) $CaCl_2$ and $BaCl_2$ (4) $MgCl_2$ and $CaCl_2$

12. Which species has dipole moment?
(1) Chair form of Cyclohexane-1,4-dione   (2) p-chloro phenol
(3) Staggered form of 1,2-dichloro ethane   (4) Trans-2-butene

13. An open container of volume 5L contains air at 27°C and 1 atm. The container is heated to T K. The amount of air expelled from the container is found to be 1.5L at – 33°C and 1atm. What is the value of temperature T K ?
(1) 400 K   (2) 720 K   (3) 450 K   (4) 480 K

14. The stability of +1 oxidation state increases in the sequence.
(1) Ga < In < Al < Tl (2) Al < Ga < In < Tl
(3) Tl < In < Ga < Al (4) In < Tl < Ga < Al

15. The most stable conformational isomer of cis 1-Bromo-2-chloro cyclohexane is
(1) Both chlorine and Bromine atoms at axial position.
(2) Bromine atom at axial position and chlorine atom at equatorial position
(3) Both chlorine and Bromine atoms at equatorial position
(4) Bromine atom at equatorial position and chlorine atom at axial position.

16. Which of the following forms a positively charged sol ?
(1) haemoglobin   (2) clay   (3) eosin   (4) charcoal

17. Phosphine is prepared by the action of :
(1) White Phosphorus and H_2SO_4 (2) White Phosphorus and NaOH
(3) White Phosphorus and H_2S (4) White Phosphorus and HNO_3

18. Structural isomers of the compound with molecular formula $C_4H_{10}O$ are -
(1) 6   (2) 7   (3) 8   (4) 9

19. Which of the following is **NOT** correct?
(1) Fertile soils are colloidal in nature in which humus acts as a protective colloid.
(2) Latex is negatively charged colloidal solution.
(3) The formation of micelles takes place only below critical micelle concentration.
(4) In chemisorption unimolecular layer forms.

20. $2SO_2\ (g) + O_2\ (g) \xrightarrow{\ 'X'\ } 2SO_3\ (g)$
In above reaction 'X' will be :
(1) $CuCl_2$ (2) V_2O_5 (3) $Pt + Rh$ (4) 'C'

**21.** Which of the following compounds can not show geometrical isomerism?

(1)   (2)   (3)   (4)
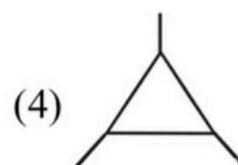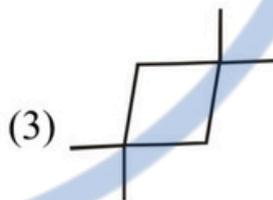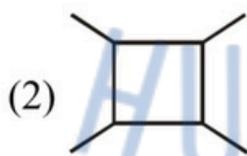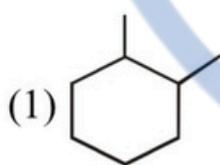
22. A sample of hard water containing Ca^{2+}(aq) ions, having concentration 400 ppm is passed through cation exchange resin (RH) which exchanges Ca^{2+} (aq) ions for H^+(aq). What is the volume of 1M NaOH required to neutralise 1L of water obtained from cation exchange resin after exchange of Ca^{2+} of hard water?

(1) 20 mL   (2) 10 mL   (3) 40 mL   (4) 80 mL

23. What will be the product when ammonia react with excess chlorine:

(1) NH_4Cl (2) N_2 (3) NCl_3 (4) $NaClO_3$

24. How many sterogenic centers and stereoisomers are present in given compound respectively?
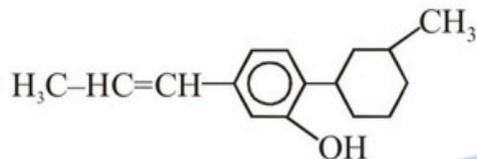
(1) 3, 8   (2) 4, 8   (3) 4, 16   (4) 3, 6

25. For a vander Waal's gas the volume of gas molecules are negligible. The molar volume of gas is 20L/mol at 300K. The compressibility factor of the gas is -
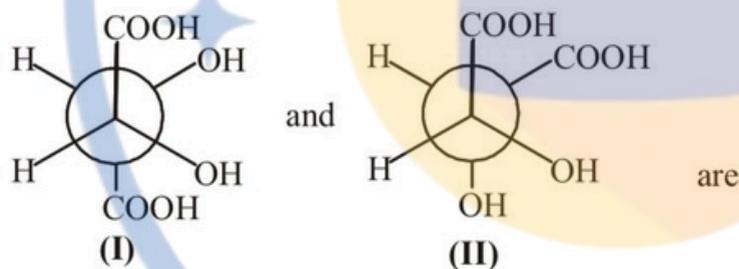
[**Given :** $a = 96$ L^2.atm.mol^{-2} and $R = 0.08$ L.atm.mol^{-1}K^{-1}]

(1) 0.8   (2) 0.75   (3) 1.2   (4) 1

26. Geometry around 'Xe' in XeF_6 is :

(1) Square bipyramidal   (2) Distorted octahedral

(3) Pentagonal pyramidal   (4) Square Antiprismatic

27. The structure :

(1) Structural isomers   (2) Enantiomer   (3) Diastereomers   (4) Identical

28. How many grams of FeC_2O_4(s) is required to react with 50 mL, 0.1 M $KMnO_4$(aq) in acidic medium?

(1) 1.2g   (2) 1.8 g   (3) 3.6 g   (4) 2.4 g

29. Which of the them is an amphoteric oxides?

(1) CaO   (2) CO_2 (3) SiO_2 (4) SnO_2

30. Choose the correct statement among the given :

(1) Gauche form of ethane-1,2-diol is most stable

(2) Enol form of Ethanamide can show geometrical

(3) In methyl cyclohexane, methyl group lies at equatorial position than axial position.

(4) All are correct

Category-2

1. The vapour pressure lowering of a 0.1 M aqueous solution of non – electrolyte at $75°C$ is $\left(\Delta H_v^0 \text{ of water}=9720\,\text{cal/mol and antilog }0.407=2.55\right)$

   (A) 0.54 torr   (B) 0.78 torr
   (C) 0.02 torr   (D) 0.22 torr

2. The C – C single bond distance is $1.54\,\overset{\text{o}}{\text{A}}$. The minimum distance between the terminal cartons in propane is $\left(\sin 54°45'=0.8136\right)$

   (A) Greater than $3.08\,\overset{\text{o}}{\text{A}}$ (B) Equal to $3.08\,\overset{\text{o}}{\text{A}}$

   (C) $2.51\,\overset{\text{o}}{\text{A}}$ (D) Can't be calculated

3. A sample of aqueous solution of acetic acid $\left(0.5\,\text{mol CH}_3\text{COOH in }10\,\text{mol of H}_2\text{O}\right)$ is mixed with a sample of aqueous solution of $\text{NaOH}\left(0.5\,\text{mol NaOH in }10\,\text{mol of H}_2\text{O}\right)$. The freezing point of the solution is ($K_{f,\ H2O}$=1.86)

   (A) -5°C   (B) -5.16°C
   (C) 0°C   (D) 5.16°C

4. Hydrogen peroxide can be prepared by the successive reactions
   $$2NH_4HSO_4 \rightarrow \left(NH_4\right)_2 S_2O_8$$
   $$\left(NH_4\right)_2 S_2O_8 + 2H_2O \rightarrow 2NH_4HSO_4 + H_2O_2$$
   The first is an electrolytic reaction, and the second a steam distillation.
   What current would have to be used in the first reaction to produce enough intermediate to yied 100 g pure H_2O_2 per hour? Assume 50% current efficiency.

   (A) 158 A   (B) 315 A
   (C) 102 A   (D) 204 A

5. Equilibrium constant for the dissociation of $Ag\left(NH_2\right)_2^+$ into Ag^+ and NH_3 is 6×10^{-8}. Calculate $E°$ for the following half reaction, $Ag\left(NH_3\right)_2^+ + e \rightarrow Ag(s) + 2NH_3$

   $\left(E°_{Ag^+/Ag} = 0.799V, \log 6 = 0.7781\right)$

   (A) 0.37   (B) 0.799
   (C) 0.7781   (D) 0.6

Category-3

6. A student is analysing a salt in the laboratory. When he added dilute HCl to the salt solution, a gas is evolved, when the gas is exposed to a filter paper moistened with potassium iodate and starch solution, he observed blue colouration of the filter paper. Then he added $BaCl_2$ solution to the salt solution and observed white precipitate immediately. He found that the precipitate is soluble in dilute mineral acids. The anion of the salt is

(A) $S_2O_3^{2-}$

(B) SO_3^{2-}

(C) CO_3^{2-}

(D) HCO_3^-

7. Which of the following complexes is/are chiral?

(A) $cis-[Pt(NH_3)_2 Cl_2]$

(B) $[Zn(H_2O)_2 Cl_2](\text{tetrahedral})$

(C) $\left[Pt(C_2O_4)_3\right]^{4-}$

(D) $trans - \left[Pt(NH_3)_4 BrCl\right]^2$

8. For the reaction $Cu^{2+} + 2e \rightarrow Cu$

$\log\left[Cu^{2+}\right]$ vs. E_{red} graph can be plotted as following (At NTP)
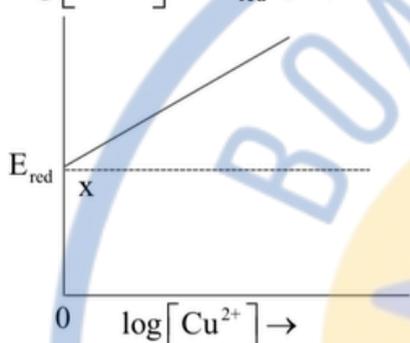
Where OX = 0.34 V then electrode potential of the half cell of $Cu/Cu^{2+}(0.1M)$ will be:

(A) 0.3695 v

(B) $-$ 0.3105 V

(C) 0.3991 V

(D) -0.2809 V

9. An ideal gas expanding isothermally against the external pressure equal to zero, shows

(A) $\Delta E = 0$

(B) $W = 0$

(C) $q = 0$

(D) $\Delta H = +ve$

10. Electrolysis of molten $MgCl_2$ is the final production step in the isolation of magnesium from sea water by Dow's process. Assuming that 35.6 g of Mg metal forms. Which of the following statement (s) is/are correct?

(A) 2.96 moles of electrons are required

(B) $3.83 \times 10^5 C$ charge is required

(C) 31.8 A of current is required for 2.5 hours

(D) 41.5 A of current is required for approximately 115 mins